Bonding Practice Test

Using Lewis structures: Draw out or model how an ionic or covalent substance is created.
- We demonstrated this process in class.

Aluminum chloride **N_2 (Nitrogen gas)**

1)	Type oB bond **(I or C)**	Electrons: shared or transferred?	2) Write the compound formula formed between the two ions listed.	
Carbon tetrachloride	C	Shared	**Al^{+3} Cl^{-1}**	$AlCl_3$
NH_4ClO	I	Transfer	**Ca^{+2} OH^{-1}**	$Ca(OH)_2$
ClO_2	C	Shared	**Al^{+3} PO_3^{-3}**	$AlPO_3$
Sodium hydroxide	I	Transfer	**Sn^{+4} S^{-2}**	SnS_2

3) Compound Name	Compound Formula
Cobalt (III) chloride	$Co^{+3}Cl^{-1} \rightarrow CoCl_3$
Aluminum hydroxide $Al^{+3}OH^{-1}$	$\rightarrow Al(OH)_3$
Disilicion hexabromide	Si_2Br_6
Lead (IV) sulfate Pb^{+4} SO_4^{-2}	$Pb(SO_4)_2$
Ammonium sulfide	$(NH_4)_2S$
Calcium acetate	$Ca(C_2H_3O_2)_2$
Manganese (II) Chromate	**$MnCrO_4$**
Nitrogen tri flouride	**NF_3**
trichromium di nitride	**Cr_3N_2**
~~tri~~ Cadmium Phosphide	**Cd_3P_2**
Zinc hydroxide	**$Zn(OH)_2$**
water	**H_2O**

In the beakers below, model the dissolving of Aluminum Chloride and Sulfur dioxide in each beaker.